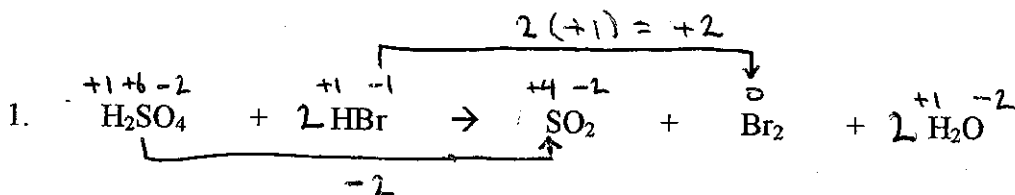
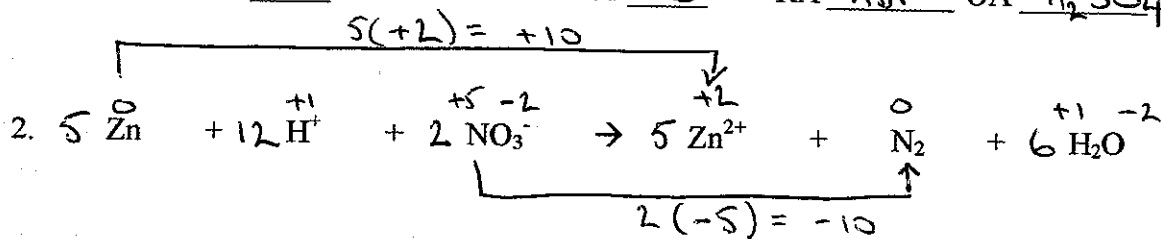


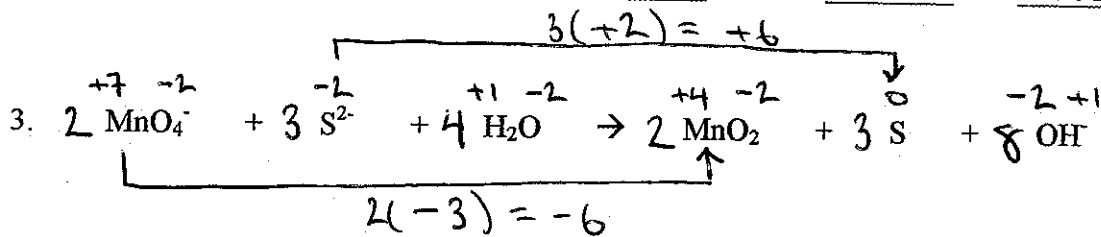
Balance the following redox reaction using the oxidation number method. Identify the oxidizing agent (OA), the reducing agent (RA), the element oxidized, and the element reduced.



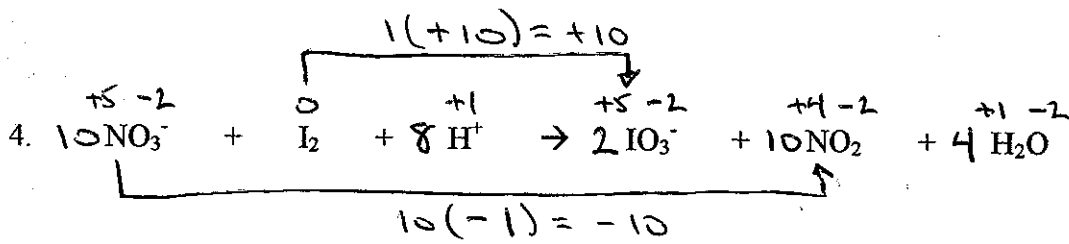
Element Oxidized Br Element Reduced S RA HBr OA H₂SO₄



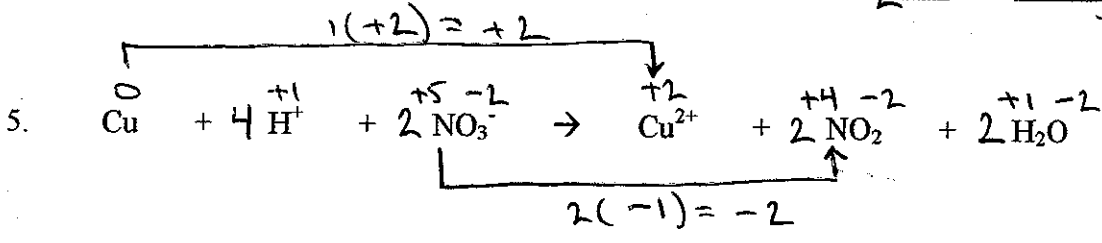
Element Oxidized Zn Element Reduced N RA Zn OA NO₃⁻



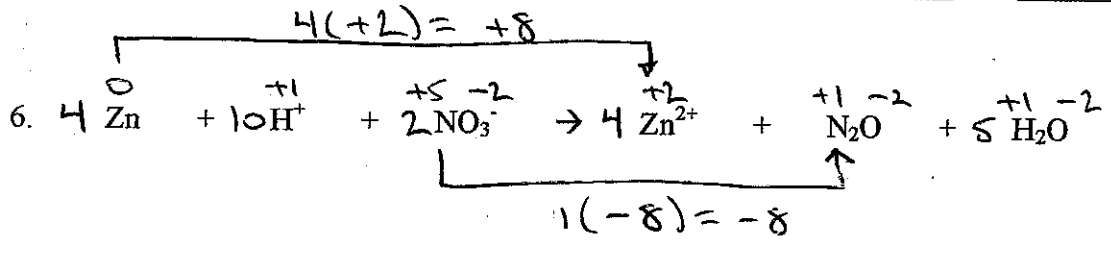
Element Oxidized S Element Reduced Mn RA S²⁻ OA MnO₄⁻



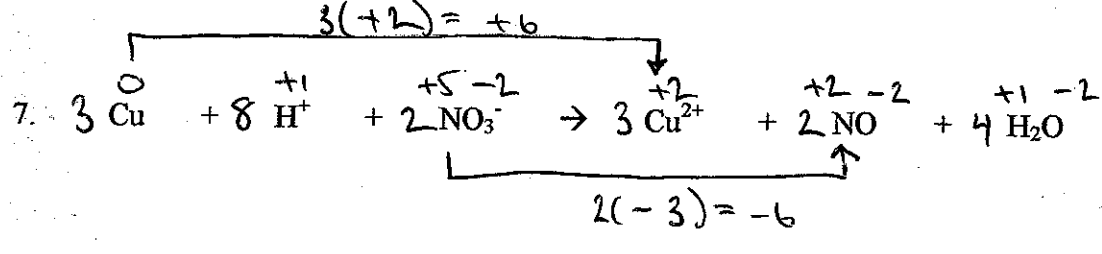
Element Oxidized I Element Reduced N RA I₂ OA NO₃⁻



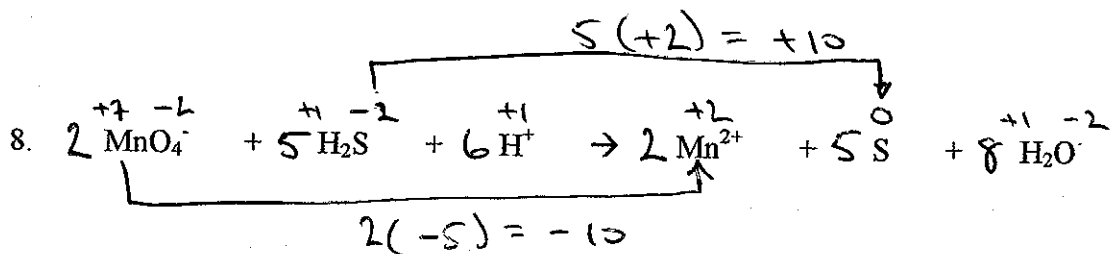
Element Oxidized Cu Element Reduced N RA Cu OA NO₃⁻



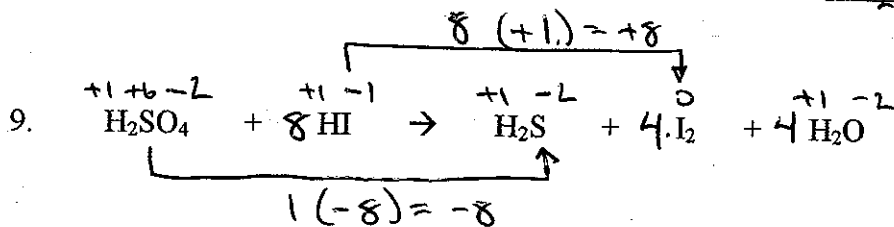
Element Oxidized Zn Element Reduced N RA Zn OA NO₃⁻



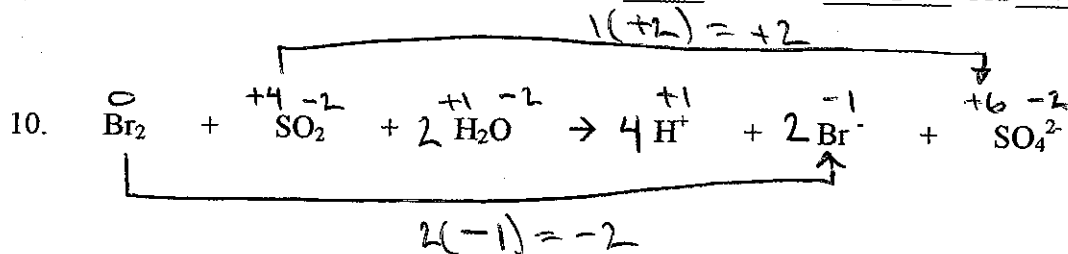
Element Oxidized Cu Element Reduced N RA Cu OA NO₃⁻



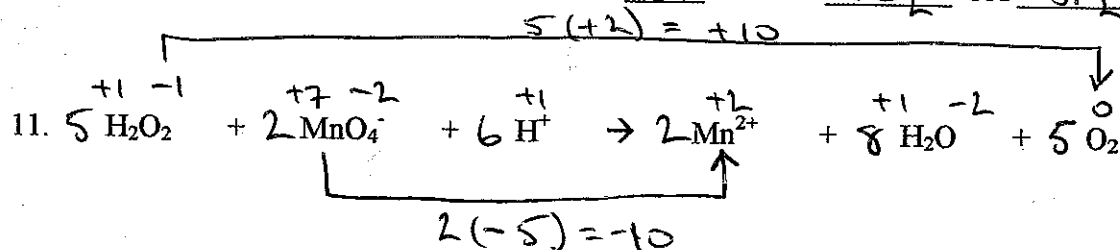
Element Oxidized S Element Reduced Mn RA H₂S OA MnO₄⁻



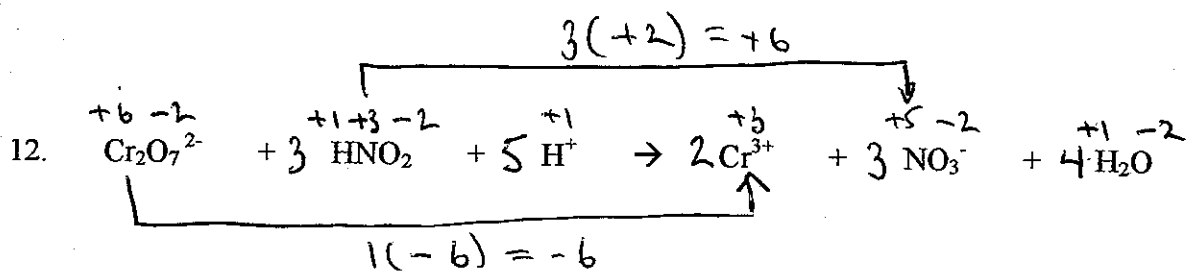
Element Oxidized I Element Reduced S RA HI OA H₂SO₄



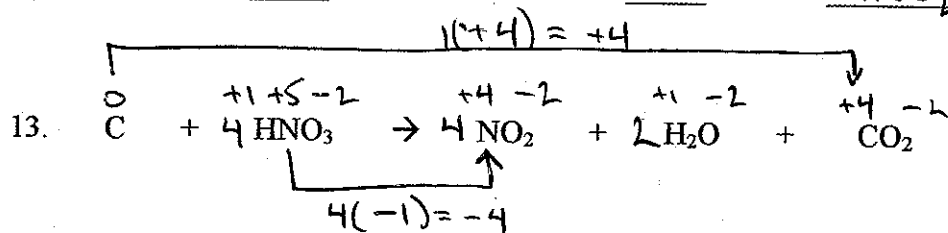
Element Oxidized S Element Reduced Br RA SO₂ OA Br₂



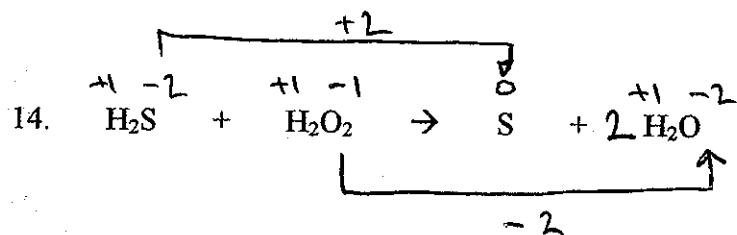
Element Oxidized O Element Reduced Mn RA H₂O₂ OA MnO₄⁻



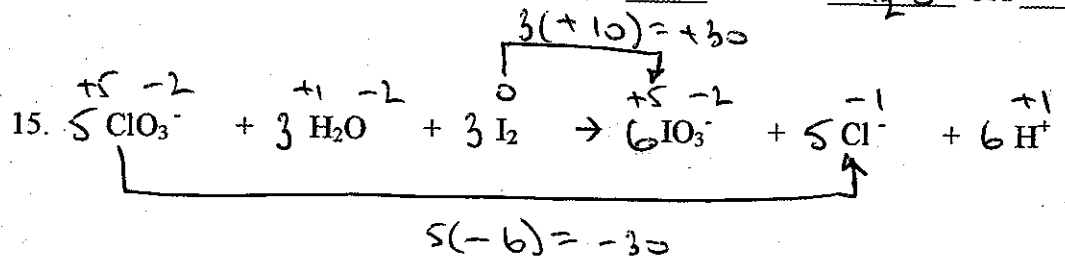
Element Oxidized N Element Reduced Cr RA HNO₂ OA Cr₂O₇²⁻



Element Oxidized C Element Reduced N RA C OA HNO₃



Element Oxidized S Element Reduced O RA H₂S OA H₂O₂



Element Oxidized I Element Reduced Cl RA I₂ OA ClO₃⁻